

### Atomic Structure Chapter 4

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#### STRUCTURE OF ATOM - 9TH CBSE ||PART 1 BY SANJIV SIR

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chapter 4: the structure of the atom Although the ancient Greeks had many creative ideas about matter, our current understanding is based upon John Dalton's atomic theory, published in 1803. Essentially, matter is made up of atoms, which are indivisible and indestructible.

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4.05 Atomic Number, Mass Number, Isotopes and Isobars 4.05 Atomic Number, Mass number, Isotopes and Isobars. Chapter 4 Structure of Atoms - Test. 05 The Fundamental Unit of Life 5 Topics | 6 Quizzes 5.01 Introduction to cell 5.01 Introduction to Cell. 10.03 Plasma Membrane 5.02 Plasma Membrane ...

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1. All elements are composed of tiny invisible particles (atoms) 2. All atoms of the same element are identical. 3. Atoms of different elements can physically combine together or chemically combine in simple whole number ratios to form compounds. 4. Chemical reactions occur when atoms are separated, joined, or rearranged.

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Dalton's Atomic theory. 4 part theory about atoms. 1) All elements are composed of \_\_\_\_\_ atoms. Atoms can combine both \_\_\_\_\_ and \_\_\_\_\_. physically & chemically. 2) atoms in a given element are \_\_\_\_\_, but different from those in other elements. identical. 3) atoms are neither \_\_\_\_\_ nor \_\_\_\_\_.

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The nucleus has two fewer protons, so the atomic number decreases by two. The nucleus also has two fewer neutrons, so the mass number decreases by four. A new element is formed.

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Chapter 4 Atomic Structure Answers

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The central nucleus of the atom contains the neutrons and positively charged protons. The nucleus is surrounded by as many orbital electrons (each with a charge of  $-1$ ) as there are. protons (each with a charge of  $+1$ ) in the nucleus. The three sub-atomic particles differ from each other in two important ways.

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